1. Indicate whether each of the following 0.10 \( M \) solutions will be ACIDIC, BASIC, or NEUTRAL

A. NaHSO\(_4\)  
B. HNO\(_3\)  
C. MgCl\(_2\)  
D. NaClO  
E. CH\(_3\)NH\(_2\)  
F. MgSO\(_3\)  

2. What is the pH of a 0.80 \( M \) solution of sodium formate (Na\(^+\)HCO\(_2^-\))?

3. What is the pH of a 0.15 \( M \) solution of sodium hydrogen phosphate, Na\(_2\)HPO\(_4\)?

4. What is the pH and the concentration of all selenium-containing species for a 0.25 \( M \) solution of selenous acid, H\(_2\)SeO\(_3\)?

\[ K_{a1} = 2.7 \times 10^{-3} \quad K_{a2} = 2.5 \times 10^{-7} \]